Stoichiometry Lab Answers

1. Cu and AgNO₃

2. There isn't a significant pattern in the mass data, although answers for this question may vary.

3. AgNO₃ to Ag should be same in each case
   - Cu\text{used} to AgNO₃ should be twice as much AgNO₃
   - Cu\text{used} to Ag should be twice as much Ag
   - Cu\text{given} to AgNO₃ no pattern

4. Patterns reflect the ratios that the balanced equation will have. There is no pattern for Copper given because it is the excess reagent.

5. Patterns for mole ratio data will be similar to patterns for mole data...but patterns should be more clear using the ratios.

6. Similar answer to number 4...the ratios can be used to determine the balanced equation.

7. Cu + 2 AgNO₃ \rightarrow Cu(NO₃)₂ + 2 Ag

8. More AgNO₃ would need to be added because it is the limiting reagent. In each of the reactions, there was leftover copper wire.